Titration practice problems:

-Write and balance equation

-Solve for moles

-multiply by mole ratio

-solve for volume or molarity of unknown

1. Sulfuric acid is titrated against 25.00cm3 of 0.2000 mol dm-3 NaOH. It is found that 23.20cm3 of sulfuric acid is required for neutralization. Calculate the concentration of the sulfuric acid.
2. Making a standard solution (reference page30) What mass of sodium sulfate must be used to make up 250cm3 of a 0.100 mol dm-3 solution?
   1. What is the concentration of sodium ions in solution?
3. Calcium carbonate is reacted with 50.0 cm3 of 0.500 mol dm-3 hydrochloric acid. What mass of calcium carbonate is required for an exact reaction?
4. For neutralization, 25.00cm3 of phosphoric acid requires 28.70cm3 of sodium hydroxide of concentration 0.1500 mol dm-3. What is the concentration of the phosphoric acid?

Also:

Page 30 Question 28, 29

Page 32 Question 31