**Snacks Lab Follow-up**

**Calculations**:

Complete for each type of nut: (it may be helpful to put the values in a data table)

1. Calculate the heat of combustion of each nut. (units should be in J/g, so remember to divide by the mass of the nut)
2. Convert your answer to Cal/g so you can compare the values to the known values for the nut (from the package).
3. Calculate your % error

**((valueexp - valueaccepted)/(valueaccepted))x100 = % error**

**Analysis**:

1. Use yours.

**Questions**:

1. What mass of water could be heated by burning 2g of paraffin (wax) from 0 °C to 25 °C? (heat of combustion of combustion of paraffin is 10.0 Kcal/g)
2. What would the final temperature of a 5L sample of water be, beginning at 0°C, if 3 g of propane were burned to heat the water? (the heat of combustion of propane is 50.0 kJ/g)

**Turn in Lab Write-up on a separate piece of paper.**

**Remember Write-up should include:**

**Pre-lab**

**Purpose**

**Procedure**

**Data (in a table)**

**Calculations**

**Analysis**

**Questions**