

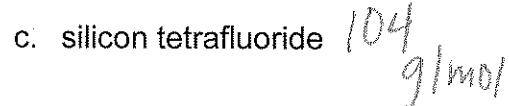
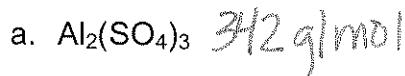
## CHAPTER 7 REVIEW

*Baker*

Write a chemical formula for each substance.

- |                                   |                      |   |
|-----------------------------------|----------------------|---|
| 1. <u><chem>NaCl</chem></u>       | sodium chloride      | Name each of the following substances:                        |
| 2. <u><chem>N2O5</chem></u>       | dinitrogen pentoxide | 26. <u>gold (III) bromide</u> <chem>AuBr3</chem>              |
| 5. <u><chem>H2SO3</chem></u>      | sulfurous acid       | 27. <u>carbon monoxide</u> <chem>CO</chem>                    |
| 6. <u><chem>AgC2H3O2</chem></u>   | silver acetate       | 28. <u>potassium carbonate</u> <chem>K2CO3</chem>             |
| 8. <u><chem>SO3</chem></u>        | sulfur trioxide      | 29. <u>chlorous acid</u> <chem>HClO2</chem>                   |
| 9. <u><chem>N2O4</chem></u>       | dinitrogen tetroxide | 31. <u>nickel permanganate</u> <chem>Ni(MnO4)2</chem>         |
| 10. <u><chem>Ca(OH)2</chem></u>   | calcium hydroxide    | 32. <u>aluminum sulfate</u> <chem>Al2(SO4)3</chem>            |
| 11. <u><chem>Li2S</chem></u>      | lithium sulfide      | 33. <u>carbonic acid</u> <chem>H2CO3</chem>                   |
| 12. <u><chem>H3PO3</chem></u>     | phosphorous acid     | 34. <u>barium acetate</u> <chem>Ba(C2H3O2)2</chem>            |
| 13. <u><chem>N2O</chem></u>       | dinitrogen monoxide  | 35. <u>manganese(II) hydroxide</u> <chem>Mn(OH)3</chem>       |
| 14. <u><chem>Fe2O3</chem></u>     | Iron (III) oxide     | 36. <u>potassium dihydrogen phosphate</u> <chem>KH2PO4</chem> |
| 15. <u><chem>PbClO3</chem></u>    | lead(II) chlorate    | 38. <u>aluminum bromate</u> <chem>Al(BrO3)3</chem>            |
| 17. <u><chem>ZnCl2</chem></u>     | zinc chloride        | 40. <u>sulfur trioxide</u> <chem>SO3</chem>                   |
| 18. <u><chem>Ca3(PO4)2</chem></u> | calcium phosphate    | 41. <u>potassium cyanide</u> <chem>KCN</chem>                 |
| 19. <u><chem>OF2</chem></u>       | oxygen difluoride    | 42. <u>lead nitrate</u> <chem>Pb(NO3)2</chem>                 |
| 20. <u><chem>NaNO2</chem></u>     | sodium nitrite       | 44. <u>iron (III) chloride</u> <chem>FeCl3</chem>             |
| 21. <u><chem>Fe2(SO4)3</chem></u> | iron(III) sulfate    | 45. <u>sulfur hexafluoride</u> <chem>SF6</chem>               |
| 22. <u><chem>HBr</chem></u>       | hydrobromic acid     | 46. <u>calcium nitride</u> <chem>Ca3N2</chem>                 |
| 23. <u><chem>Cr2O3</chem></u>     | chromium(III) oxide  | 47. <u>copper(I) iodide</u> <chem>CuI</chem>                  |
| 24. <u><chem>(NH4)2CO3</chem></u> | ammonium carbonate   | 48. <u>silicon dioxide</u> <chem>SiO2</chem>                  |
|                                   |                      | 49. <u>tin (IV) acetate</u> <chem>Sn(C2H3O2)4</chem>          |

1. Calculate the molar mass:



2. How many moles are contained in 18.0 grams of NaCl?

$$0.3 \text{ mol NaCl}$$

3. What is the mass in grams of 0.33 moles of  $\text{CuSO}_4$ ?

$$53 \text{ g CuSO}_4$$

4. What is the mass in grams of  $4.7 \times 10^{22}$  molecules of CaS?

$$5.6 \text{ g CaS}$$

5. Calculate the number of moles contained in 25.0 grams of iron (II) chloride.

$$0.2 \text{ mol FeCl}_2$$

6. Calculate the percent composition for the following compounds:



c. aluminum hypochlorite

